1	(a	(i)	a compound which contains carbon and hydrogen only	[1]
		(ii)	alkanes contain <b>only</b> C-C single bonds <b>or</b> they are saturated (hydrocarbons) <b>or</b> have the general formula C <sub>n</sub> H <sub>2n+2</sub>	[1]
			alkenes contain at least one C=C double bond <b>or</b> they are unsaturated (hydrocarbons) <b>or</b> have the general formula $C_nH_{2n}$	[1]
	(b)	C <sub>20</sub>	$H_{42} \rightarrow 2C_4H_8 + 2C_2H_4 + C_8H_{18}$	[1]
	(c)	(i)	any unambiguous structure of BrCH <sub>2</sub> CH <sub>2</sub> Br <b>NOT</b> just C <sub>2</sub> H <sub>4</sub> Br <sub>2</sub>	[1]
		(ii)	CH <sub>3</sub> -CH=CH-CH <sub>3</sub> For any butene [1] only	[2]
		(iii)	(CH <sub>3</sub> -CH <sub>2</sub> -CH=CH <sub>2</sub> ) + H <sub>2</sub> O [1] $\rightarrow$ CH <sub>3</sub> -CH <sub>2</sub> -CH <sub>2</sub> -CH <sub>2</sub> OH [1] ALLOW CH <sub>3</sub> -CHOH-CH <sub>2</sub> -CH <sub>3</sub> butene reacts with water/steam (to form butanol) ONLY [1]	[2]
		(iv)	$C_6H_{12} + H_2 \rightarrow C_6H_{14}$ alkenes react with <b>hydrogen</b> [1] <b>ONLY</b>	[2]
	(d)	vol	ume of oxygen used = $150  \text{cm}^3$	[1]
volume of carbon dioxide formed = $100 \text{ cm}^3$ any equation of the combustion of an alkene			[1]	
		form	$2C_5H_{10} + 15O_2 \rightarrow 10CO_2 + 10H_2O$ nulae <b>ND</b> balancing	[1] [1]

2	(a	(i)	amino acid / peptides; salt / carboxylate or soap / fatty acid or glycerine / alcohol; sugars or glucose; <b>accept:</b> named sugar	[1] [1] [1]
		(ii)	polyester; allow: named polyester polyamide; allow: nylon	[1] [1]
	(b)	<ul> <li>b) one correct amide linkage; second amide linkage correctly orientated</li> <li>– NHCO – followed by – NHCO –</li> <li>note: monomers are amino acids not diamines or dicarboxylic acid</li> </ul>		[1] [1]
	(c)	uns sat <b>or:</b>	<ul> <li>mine / bromine water / aqueous bromine;</li> <li>saturated - brown / orange to colourless not: clear</li> <li>urated - stays brown / orange</li> <li>alkaline potassium manganate(VII);</li> <li>from purple / pink to green / brown;</li> <li>stays purple;</li> <li>acidic potassium manganate(VII)</li> <li>from purple / pink to colourless; not: clear</li> </ul>	[1] [1] [1]
			stays purple;	[Total: 10]

(i)	Zn + 2HC $l \rightarrow$ ZnC $l_2$ + H <sub>2</sub> not balanced = [1]	[2]
(ii)	3 bps and 1 nbp around As; 1 bp each hydrogen atom;	[1] [1]
	(97.4/75 =) 1.3 <b>and</b> (2.6/1 = ) 2.6; empirical formula AsH <sub>2</sub> ; <b>note:</b> correct formula with no working = [1]	[1] [1
(ii)	As <sub>2</sub> H <sub>4</sub> ;	[1]
(iii)	$H_2As-AsH_2/AsH_2-AsH_2;$	[1
(	amide / peptide;	[1]
(ii)	named strong acid / alkali; <b>allow:</b> HC1/ enzymes	[1]
(iii)	amino acid; allow: peptides	[1]
(	Cu and As have more than one oxidation state / valency;	[1]
(ii)		[2]
		[Total: 14]
	(ii) (ii) (iii) ( (iii) (iii) (	<ul> <li>(ii) 3 bps and 1 nbp around As; 1 bp each hydrogen atom;</li> <li>(97.4/75 =) 1.3 and (2.6/1 = ) 2.6; empirical formula AsH<sub>2</sub>; note: correct formula with no working = [1]</li> <li>(ii) As<sub>2</sub>H<sub>4</sub>;</li> <li>(iii) H<sub>2</sub>As-AsH<sub>2</sub>/AsH<sub>2</sub>-AsH<sub>2</sub>;</li> <li>( amide / peptide;</li> <li>(ii) named strong acid / alkali; allow: HCl / enzymes</li> <li>(iii) amino acid; allow: peptides</li> </ul>

4	(a)	(i)	correct structural formula of ethanoic acid allow: –OH not: –COOH	[1]
		(ii)	correct structural formula of ethanol allow: –OH	[1]
	(b)	(i)	ethyl ethanoate	[1]
		(ii)	-OC <sub>6</sub> H <sub>4</sub> COOCH <sub>2</sub> CH <sub>2</sub> O- correct ester linkage correct repeat units continuation <b>accept:</b> boxes if it is clear what the box represents	[1] [1] [1]
		(iii)	any <b>two</b> from: long time to decay landfill sites visual pollution / litter danger to animals poisonous gases when burnt <b>accept:</b> any correct suggestion	[2]
	(c)	pro	thetic – only two monomers tein – many different monomers	[1] [1]
		nylo	tein has 1 C=O and 1N–H on has  2 C=O / 2N–H	[1] [1]
		<b>or:</b> synthetic – one monomer is a dicarboxylic acid and the other is a diamine protein all monomers are amino acids		

5	(a	(i)	ethanol CH <sub>3</sub> -CH <sub>2</sub> -OH	[1] [1]
			propanoic acid $CH_3$ - $CH_2$ - $COOH$ independent marking, no ecf <b>accept</b> $C_2H_5$ <b>not</b> – HO	[1] [1]
		(ii)	type of compound – salt / sodium carboxylate / alkanoate <b>not</b> soap / sodium stearate etc use – soap / cleaning / detergent	[1] [1]
		(iii)	terylene / PET / Dacron / diolen / mylar / crimplene	
	(b)		polyamide / amide / peptide / polypeptide	[1]
		(ii)	correct amide linkage <u>NHCO then CONH</u> <b>cond</b> to mark 1, 2 monomers (different shading in box) <b>cond</b> continuation (to <b>ONE</b> correct linkage)	[1] [1] [1]
			<b>OR</b> nylon 6 only one linkage – NHCO <b>cond</b> only one monomer <b>cond</b> continuation (to correct linkage)	[1] [1] [1]
		(iii)	use locating agent measure distance travelled by sample / travelled by solvent front <b>cond</b> this is $R_f = 0.5$ for mark 3, either mark 1 or mark 2 must be awarded	[1] [1] [1]
			<b>accept</b> run a chromatogram of glycine [1] compare with sample same position [1] max [2]	